



UNIVERSITY OF MYSORE

Postgraduate Entrance Examination June/July 2017

SUBJECT CODE : **45**

QUESTION BOOKLET NO.

Entrance Reg. No.

115262

QUESTION BOOKLET

(Read carefully the instructions given in the Question Booklet)

COURSE : **M.Sc.**

SUBJECT : **Chemistry**

MAXIMUM MARKS : 50

MAXIMUM TIME : ONE HOUR

(Including initial 10 minutes for filling O.M.R. Answer sheet)

INSTRUCTIONS TO THE CANDIDATES

1. The sealed questions booklet containing 50 questions enclosed with O.M.R. Answer Sheet is given to you.
2. Verify whether the given question booklet is of the same subject which you have opted for examination.
3. Open the question paper seal carefully and take out the enclosed O.M.R. Answer Sheet outside the question booklet and fill up the general information in the O.M.R. Answer sheet. If you fail to fill up the details in the form of alphabet and signs as instructed, you will be personally responsible for consequences arising during scoring of your Answer Sheet.
4. During the examination:
 - a) Read each question carefully.
 - b) Determine the Most appropriate/correct answer from the four available choices given under each question.
 - c) Completely darken the relevant circle against the Question in the O.M.R. Answer Sheet. For example, in the question paper if "C" is correct answer for Question No.8, then darken against Sl. No.8 of O.M.R. Answer Sheet using Blue/Black Ball Point Pen as follows:

Question No. 8. (A) (B) (C) (D) (Only example) (Use Ball Pen only)

5. Rough work should be done only on the blank space provided in the Question Booklet. Rough work should not be done on the O.M.R. Answer Sheet.
6. If more than one circle is darkened for a given question, such answer is treated as wrong and no mark will be given. See the example in the O.M.R. Sheet.
7. The candidate and the Room Supervisor should sign in the O.M.R. Sheet at the specified place.
8. Candidate should return the original O.M.R. Answer Sheet and the university copy to the Room Supervisor after the examination.
9. Candidate can carry the question booklet and the candidate copy of the O.M.R. Sheet.
10. The calculator, pager and mobile phone are not allowed inside the examination hall.
11. **If a candidate is found committing malpractice, such a candidate shall not be considered for admission to the course and action against such candidate will be taken as per rules.**

INSTRUCTIONS TO FILL UP THE O.M.R. SHEET

1. There is only one most appropriate/correct answer for each question.
2. For each question, only one circle must be darkened with BLUE or BLACK ball point pen only. Do not try to alter it.
3. Circle should be darkened completely so that the alphabet inside it is not visible.
4. Do not make any stray marks on O.M.R. Sheet.

ಗಮನಿಸಿ : ಸೂಚನೆಗಳ ಕನ್ನಡ ಆವೃತ್ತಿಯು ಈ ಪುಸ್ತಕದ ಹಿಂಭಾಗದಲ್ಲಿ ಮುದ್ರಿಸಲ್ಪಟ್ಟಿದೆ.

SEAL

SEAL

- 1) Buckminster fullerene C-60 is a new allotrope of carbon and is considered to be most symmetric molecule known, which is build from
- (A) Six member and five member carbon rings involving sp^2 hybridization
 - (B) Seven member and five member carbon rings involving sp^3 hybridization
 - (C) Five member and four member carbon rings involving sp^3 hybridization
 - (D) None of the above but only due to delocalization of electrons
- 2) In the quantitative analysis of nitrogen by Kjeldhal method initially the compound is digested with concentrated sulphuric acid and similarly in case of sulphur determination by Carius method the compound is treated with either peroxy or nitric acid. The formed species in those reactions respectively are
- (A) Ammonia and sulphite
 - (B) Ammonium sulphate and sulphate
 - (C) Ammonium sulphate and thiosulphate
 - (D) Ammonium sulphate and sulphide
- 3) In benzene ring pi-electrons are loosely held therefore it prefers to undergo
- (A) Nucleophilic substitution
 - (B) Free radical substitution
 - (C) Electrophilic substitution
 - (D) Complexation reaction
- 4) When toluene is treated first with oxidant followed by nitric acid and sulphuric acid and in another instant the same compound is treated with reagents in reverse order. The products formed are
- (A) ortho-nitrobenzoic acid and the mixture of ortho-and para nitro benzoic acids
 - (B) para-nitrobenzoic acid and the mixture of ortho-and meta-nitro benzoic acids
 - (C) para-nitrobenzoic acid and meta nitro benzoic acid
 - (D) Meta-nitrobenzoic acid and the mixture of ortho-and para nitro benzoic acids

- 5) When ethyl benzene is treated with chlorine gas under light source or heat, it produces 1-phenyl-1-chloro ethane with 91% yield and 2-phenyl-chloro ethane with only 9% yield
- (A) Since hydrogen bonded to second carbon of benzylic group.
(B) Methyl group of ethane group is withdrawing the electrons
(C) Since benzylic hydrogen can easily abstracted by the halide
(D) None of the above
- 6) When triphenyl chloro methane is treated with zinc in benzene at about 180 degree it produces hexaphenyl ethane. This product formation could be explained via
- (A) Substitution mechanism (B) Redox reaction
(C) Free radical mechanism (D) None of the above
- 7) Carbon dioxide is an environmental concern gas but it can be sensibly used for industrial purpose. When it is treated with Grignard reagent followed by an acid it produces very useful material
- (A) an alcohol (B) a phenol
(C) a dicarboxylic acid (D) a carboxylic acid
- 8) When meta- toluic acid is treated with lithium aluminum hydride it produces
- (A) m- methyl phenol (B) m-methyl chlorobenzene
(C) m-methyl benzyl alcohol (D) None of the above
- 9) Para nitro benzoic acid is a stronger acid than benzoic acid. It is not only favored by substituent but also enthalpy, entropy and obviously Gibb's free energy. In this case it is decided by substituent which is having the nature
- (A) Election donating (B) Electron withdrawing
(C) Non polarisable (D) Resonance structure destabilizing
- 10) When copper and iron are corroded the products formed respectively are
- (A) Copper carbonate and iron carbonate
(B) Copper hydroxide and iron hydroxide
(C) Copper oxide and iron oxide
(D) Copper (ii) hydroxo carbonate and hydrated ferric oxide

- 11) When commercial polymer is manufactured it normally produces the same polymer with different degrees of polymerization. It is therefore, number average molecular weight and weight average molecular weight are determined in such cases they do not agree one another indicating that
- (A) Polymer is a heterogeneous mixture (B) Single polymer
(C) Less homogeneous (D) None of the above
- 12) Silicone polymers $\text{HO}(\text{SiR}_2\text{O})_n\text{H}$ are finding large applications due to their stability light weight and bio-acceptability in addition they retain their structure in the temperature
- (A) Normal temperature to above boiling point temperature
(B) Zero degree to 200 degree
(C) - 90 degree to about 150 degree
(D) Zero degree Kelvin to 300 degree
- 13) Ziegler- Natta catalyst is
- (A) a mixture of beryllium titanium salts
(B) a mixture alkyl aluminium chloride and cobaltous chloride
(C) a mixture of di ethyl aluminium chloride and titanium chloride
(D) alkyl titanium chloride
- 14) Fuel cell involves combustible fuel and it is oxidized at anode and the oxidant is reduced at the cathode. Therefore in the hydrogen oxygen fuel cell the net cell reaction must involve
- (A) Hydrogen and carbon dioxide (B) Hydrogen and carbon monoxide
(C) Hydrogen and nitric oxide (D) Hydrogen and oxygen
- 15) Nickel cadmium cell is involving cadmium and nickel in presence of alkali in the cadmium and nickel change their oxidation states respectively from
- (A) +3 to +2 and 0 to +2 (B) +4 to +3 and 0 to +2
(C) +2 to +4 and 0 to +0 (D) None of the above
- 16) Among the allotropic forms of sulphur with 8-atoms and sulphur with 6-atoms are normally stable and stability of those can be attributed to the respective conformation
- (A) Octane and benzene like
(B) Octane and hexane like
(C) Puckered cyclohexane and twisted octane like
(D) Crown and chair like

- 17) When a beam of x-rays of definite wavelength is falling on the face of crystal, the successive layers of atom presenting planes parallel to the crystal face and are equi distance on one another and the reflect the rays which will reinforce only when the path difference is
- (A) $\lambda = d \sin \theta$ (B) $\lambda = d/2 \sin \theta$
(C) $2\lambda = 3d \sin \theta$ (D) $n\lambda = 2d \sin \theta$
- 18) By rotating the mounted crystal the angle of incidence is varied the reflection from the planes (100) of the rock salts the maximum intensity are found to be at 5.9 degree 11.85 degree and 18.15 degree, the sign angles of these are found to be 0.103, 0.250 and 0.312. These are in the approximate ratio and account for Bragg's equation.
- (A) 2 : 1 : 3 (B) 3 : 1 : 3
(C) 1 : 2 : 3 (D) None of the above
- 19) From the study of dipole moment and degree of hydrogen bond formation of certain molecules it would be possible to distinguish between their possible geometries, for example, maleic acid and fumaric acid have none positive dipole moment, therefore, their trans isomers must have
- (A) Dipole moment values greater than cis isomer
(B) Negative dipole moment greater than cis isomer
(C) Zero dipole moment
(D) None of the above
- 20) When an aldehyde is treated with tri ethyl borane followed by an acid the products obtained are
- (A) Ethane and phenol (B) Ethane and carboxylic acid
(C) Ethane and alcoho (D) Ethene and alcohol
- 21) Gold and silver can be extracted from their finely powdered ores initially by treating with aqueous sodium cyanide followed by zinc. The respective reactions involved are the formation of their
- (A) Cyanides and oxidation
(B) Metals
(C) Cyano complexès and reduction to metals
(D) None of the above

- 22) When the molecular weight of acetic acid in its very dilute solution of benzene is determined by cryoscopic method. The value found is about 110; it is therefore the form in which it is present in solution is
- (A) Dissociated form (B) Monomeric form
(C) Dimeric form (D) Fragmented
- 23) Carbon dioxide in normal condition is having zero dipole moment whereas similar molecules, water and sulphur dioxide are with non zero dipole moment values. These properties can be due to
- (A) No electro negativity between the combined atoms
(B) Planar structures
(C) There are no lone pair electrons
(D) Linear and angular geometries
- 24) There is a semi empirical method for calculating the lattice energies without Madelung constants of ionic compounds containing non spherical ions. Such lattice energy equation is called
- (A) Born-Landé equation (B) Born-Haber equation
(C) Landé equation (D) Kapustinskii's equation
- 25) First order NMR spectra of compounds show separate lines for each distinct magnetic nucleus. It is therefore, NMR spectrum of ethanol and F^{19} ($I = \frac{1}{2}$) spectrum of SF_4 must have the lines
- (A) Two lines and two lines (B) Four lines and three lines
(C) Two lines and one line (D) None of the above
- 26) Radioactive decay constant, λ of radio active isotope is derived from the equation leading to the equation containing half life period $\lambda = 0.693/t_{1/2}$.
- (A) $N_0 = N e^{-\lambda t}$ (B) $N = N_0 e^{-\lambda t}$
(C) $N_0 = N_0 e^{-\lambda t}$ (D) $N = N_0 e^{-\lambda t}$
- 27) General equation for the calculation of magnetic moment of a metal ion in a compound which is involving spin orbit coupling is given by the equation $\mu = g [J(j+1)]^{1/2}$ Bohr magneton where $g = 1 + j(j+1) + S(S+1) - L(L+1)/2(j+1)$. This equation is more or less followed by
- (A) transition elements (B) main block elements
(C) actinides (D) lanthanides

- 28) Methane, ammonia and water are isoelectronic but have different structure and properties; these could be due to central atom
- sp^3 hybridisation and no lone pair of electron
 - sp^3 hybridisation with different no of lone pair of electron
 - dsp^2 hybridisation and hydrogen bonding
 - none of the above
- 29) $[Ti(H_2O)_6]^{3+}$ shows maximum absorption band at 490 nanometer in its absorption spectrum. This could be assigned to the electronic transition and if water is replaced by CN^- the respective transition and position of maximum absorption can be indicated by
- e_g to t_{2g} and 490 nm
 - t_{2g} to e_g and 490 nm
 - t_{2g} to e_g and < 490 nm
 - CT and > 490
- 30) The fundamental vibrational frequency of $H^{35}Cl$ and $D^{35}Cl$ occur respectively at 2886 and 2891 cm^{-1} this can be attributed to
- difference in isotopic masses of chlorine
 - too difference in bond energies
 - difference in hydrogen bonding
 - difference in reduced masses of H and D
- 31) the relation ship between the rate constant, k of a reaction the activation energy, E and the frequency factor A and reaction temperature T is an important that could be made use for calculating thermo dynamic factors. This relationship is very sensitive to activation energy and is given by the relation
- $A = k e^{-E/RT}$
 - $E = k e^{-A/RT}$
 - $k = A e^{-E/RT}$
 - none of the above
- 32) Reverse osmosis could be made use even for the purification of water when a suitable semi permeable membrane is available and the force required for this comes from
- outside application
 - built within the system due to pressure difference
 - due to natural diffusion
 - none of the above

- 33) Decomposing of ammonia at about 1130 K on the surface of platinum catalyst follows
- (A) first order kinetics (B) second order kinetics
(C) fractional order kinetics (D) zero order kinetics
- 34) When arsenic is doped with germanium it forms the semi conductor which is of the type
- (A) p-type semi conductor (B) n-type
(C) npn-type (D) pn -type
- 35) The number of lone pair electron present on central atom of molecule, NH_3 , H_2O , NH_4^+ and H_3O^+ are in the respective order
- (A) (1, 0, 3, 2) (B) (1, 0, 2, 1)
(C) (1, 0, 1, 2) (D) (1, 2, 0, 1)
- 36) For a given system at equilibrium having coexisting different phases can be explained with variables pressure, temperature and composition. Such phases of equilibrium are given by the general equation involving components, C, Phases P and degrees of freedom, F, which is
- (A) $F + 2 = C + P$ (B) $C = F + 2 - P$
(C) $C - P + 2 = F$ (D) none of the above
- 37) CO and N_2 are isoelectronic species with same bond order but they show different properties and different stretching frequencies which could be attributed to
- (A) electron deficiency (B) variation in abundance
(C) electronegativity difference (D) none of the above
- 38) pH value of the solution having hydrogen ion concentration, 1.8×10^{-5} is
- (A) 5 (B) 4
(C) 5.1 (D) 4.7
- 39) Weak acid and strong base titration are to be buffered with a weak acid and its salt. If acetate buffer is used the relation between the pH of the solution and pKa of acetic acid are related through an equation called Hunderson-Hasselbatch-equation which is
- (A) $\text{pKa} = \text{pH} + \log \left[\frac{\text{CH}_3\text{COO}^-}{\text{CH}_3\text{COOH}} \right]$
(B) $\text{pH} = \text{pKa} - \log \left[\frac{\text{CH}_3\text{COO}^-}{\text{CH}_3\text{COOH}} \right]$
(C) $\text{pH} = \text{pKa} + \log \left[\frac{\text{CH}_3\text{COO}^-}{\text{CH}_3\text{COOH}} \right]$
(D) $\text{pH} = \text{pKa} - \log \left[\frac{\text{CH}_3\text{COOH}}{\text{CH}_3\text{COO}^-} \right]$

- 40) When chloride is determined from a sample the average of the four determinations (4) is found to be 32.46 and it is given that f_{99} , SD 0.17. The confident limit is calculated by
- (A) average value + or - $f_{99} \times SD$
 (B) average value + $f_{99} \times SD$
 (C) average value - $f_{99} \times SD$
 (D) average value + or - f_{99} / SD
- 41) CN_2 , SCN_2 , and $SeCN_2$ are derived from the respective uni-negative ions which closely imitate halide ions with respect to their reactions with metal ions. It is therefore they are conventionally called
- (A) halogens (B) halogen family ions
 (C) pseudohalogens (D) chalconides
- 42) Which one of the following represents a set of electrophiles
- (A) Br^+ and $:CCl_2$ (B) $AlCl_3$ and Cl^-
 (C) H^+ and H_2O (D) CN^- and NH_3
- 43) The percentage of constituent A in AB is found to be 22.64, 22.54, 22.61, and 22.53. These must give arithmetic mean, mean deviation and relative deviation values
- (A) 0.6, 0.4, 2.0 (B) 60, 40, 20
 (C) +0.06, 0.045, 0.2 (D) None of the above
- 44) The most reactive among the following towards sulphonation is
- (A) toluene (B) chlorobenzene
 (C) nitrobenzene (D) meta-xylene
- 45) In oxygenated hemoglobin shows O-O stretch 1106 cm^{-1} and the it is found to be diamagnetic. It is therefore, the most probable form the oxygen is bound to iron is
- (A) oxide (B) oxygen molecule
 (C) superoxide (D) peroxide
- 46) Dimidiation of cyclopentadiene is an example of a
- (A) Friedelcraft reaction (B) Chain reaction
 (C) Condensation polymerization (D) Diels-Alder reaction

- 47) Which of the following rearrangement is not acid catalyzed.
- (A) Beckman (B) Fries
(C) Hofmann (D) Pinacol – Pinacolone
- 48) In nitration, a mixture of concentrated nitric acid and sulphuric is used. In that nitronium ion formed, which is involved in the nitration of benzene and is also a slow step. It is therefore for generation of nitronium ion nitric acid must act with respect to sulphuric acid as
- (A) base (B) strong acid
(C) neutral (D) amphitonic
- 49) the most important method for preparing heavier actinides is by heavy ion bombardment accelerated in cyclotron with target materials. In such case ${}_{99}\text{Es}^{253}$ forms new element with ejection of one neutron. Similarly, ${}_{98}\text{Cf}^{252}$ with ${}_{5}\text{B}^{11}$ will form new element with ejection of six neutrons. Those new elements formed must have the atomic number and mass number,
- (A) (101, 257 and 104, 263) (B) (96, 253 and 93, 252)
(C) (101, 256 and 103, 257) (D) (99, 253-98, 252)
- 50) Viscosity of a liquid is due to the interaction between adjacent layers of molecules and it shows resistance to its flow which is governed by Poiseuille equation for the flow of known volume of liquid through a capillary tube of known radius and length, when a simplification is made through Stokes equation and keeping the experimental conditions constant for two liquids of viscosities η'^1 and η'^2 with respective densities d_1 and d_2 experience flow times t_1 and t_2 . The relation involving these parameters must be
- (A) $\eta'^1 / \eta'^2 = d_1 t_1 / d_2 t_2$
(B) $\eta'^1 d_1 / \eta'^2 d_2 = t_1 / t_2$
(C) $\eta'^1 / \eta'^2 = d_1 / d_2$
(D) None of the above



Rough Work

ಅಭ್ಯರ್ಥಿಗಳಿಗೆ ಸೂಚನೆಗಳು

1. ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯ ಜೊತೆಗೆ 50 ಪ್ರಶ್ನೆಗಳನ್ನು ಹೊಂದಿರುವ ಮೊಹರು ಮಾಡಿದ ಪ್ರಶ್ನೆ ಪುಸ್ತಕವನ್ನು ನಿಮಗೆ ನೀಡಲಾಗಿದೆ.
2. ಕೊಟ್ಟಿರುವ ಪ್ರಶ್ನೆ ಪುಸ್ತಕವು, ನೀವು ಪರೀಕ್ಷೆಗೆ ಆಯ್ಕೆ ಮಾಡಿಕೊಂಡಿರುವ ವಿಷಯಕ್ಕೆ ಸಂಬಂಧಿಸಿದ್ದೇ ಎಂಬುದನ್ನು ಪರಿಶೀಲಿಸಿರಿ.
3. ಪ್ರಶ್ನೆ ಪತ್ರಿಕೆಯ ಮೊಹರು ಜಾಗ್ರತೆಯಿಂದ ತೆರೆಯಿರಿ ಮತ್ತು ಪ್ರಶ್ನೆಪತ್ರಿಕೆಯಿಂದ ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯನ್ನು ಹೊರಗೆ ತೆಗೆದು, ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯಲ್ಲಿ ಸಾಮಾನ್ಯ ಮಾಹಿತಿಯನ್ನು ತುಂಬಿರಿ. ಕೊಟ್ಟಿರುವ ಸೂಚನೆಯಂತೆ ನೀವು ನಮೂನೆಯಲ್ಲಿನ ವಿವರಗಳನ್ನು ತುಂಬಲು ವಿಫಲರಾದರೆ, ನಿಮ್ಮ ಉತ್ತರ ಹಾಳೆಯ ಮೌಲ್ಯಮಾಪನ ಸಮಯದಲ್ಲಿ ಉಂಟಾಗುವ ಪರಿಣಾಮಗಳಿಗೆ ವೈಯಕ್ತಿಕವಾಗಿ ನೀವೇ ಜವಾಬ್ದಾರಾಗಿರುತ್ತೀರಿ.
4. ಪರೀಕ್ಷೆಯ ಸಮಯದಲ್ಲಿ:
 - a) ಪ್ರತಿಯೊಂದು ಪ್ರಶ್ನೆಯನ್ನು ಜಾಗ್ರತೆಯಿಂದ ಓದಿರಿ.
 - b) ಪ್ರತಿ ಪ್ರಶ್ನೆಯ ಕೆಳಗೆ ನೀಡಿರುವ ನಾಲ್ಕು ಲಭ್ಯ ಆಯ್ಕೆಗಳಲ್ಲಿ ಅತ್ಯಂತ ಸರಿಯಾದ/ ಸೂಕ್ತವಾದ ಉತ್ತರವನ್ನು ನಿರ್ಧರಿಸಿ.
 - c) ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯಲ್ಲಿನ ಸಂಬಂಧಿಸಿದ ಪ್ರಶ್ನೆಯ ವೃತ್ತಾಕಾರವನ್ನು ಸಂಪೂರ್ಣವಾಗಿ ತುಂಬಿರಿ. ಉದಾಹರಣೆಗೆ, ಪ್ರಶ್ನೆ ಪತ್ರಿಕೆಯಲ್ಲಿ ಪ್ರಶ್ನೆ ಸಂಖ್ಯೆ 8ಕ್ಕೆ "C" ಸರಿಯಾದ ಉತ್ತರವಾಗಿದ್ದರೆ, ನೀಲಿ/ಕಪ್ಪು ಬಾಲ್ ಪಾಯಿಂಟ್ ಪೆನ್ ಬಳಸಿ ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯ ಕ್ರಮ ಸಂಖ್ಯೆ 8ರ ಮುಂದೆ ಈ ಕೆಳಗಿನಂತೆ ತುಂಬಿರಿ:
 ಪ್ರಶ್ನೆ ಸಂಖ್ಯೆ 8. (A) (B) (C) (D) (ಉದಾಹರಣೆ ಮಾತ್ರ) (ಬಾಲ್ ಪಾಯಿಂಟ್ ಪೆನ್ ಮಾತ್ರ ಉಪಯೋಗಿಸಿ)
5. ಉತ್ತರದ ಪೂರ್ವಸಿದ್ಧತೆಯ ಬರವಣಿಗೆಯನ್ನು (ಚಿತ್ತು ಕೆಲಸ) ಪ್ರಶ್ನೆ ಪತ್ರಿಕೆಯಲ್ಲಿ ಒದಗಿಸಿದ ಖಾಲಿ ಜಾಗದಲ್ಲಿ ಮಾತ್ರವೇ ಮಾಡಬೇಕು (ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆಯಲ್ಲಿ ಮಾಡಬಾರದು).
6. ಒಂದು ನಿರ್ದಿಷ್ಟ ಪ್ರಶ್ನೆಗೆ ಒಂದಕ್ಕಿಂತ ಹೆಚ್ಚು ವೃತ್ತಾಕಾರವನ್ನು ಗುರುತಿಸಲಾಗಿದ್ದರೆ, ಅಂತಹ ಉತ್ತರವನ್ನು ತಪ್ಪು ಎಂದು ಪರಿಗಣಿಸಲಾಗುತ್ತದೆ ಮತ್ತು ಯಾವುದೇ ಅಂಕವನ್ನು ನೀಡಲಾಗುವುದಿಲ್ಲ. ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯಲ್ಲಿನ ಉದಾಹರಣೆ ನೋಡಿ.
7. ಅಭ್ಯರ್ಥಿ ಮತ್ತು ಕೊಠಡಿ ಮೇಲ್ವಿಚಾರಕರು ನಿರ್ದಿಷ್ಟಪಡಿಸಿದ ಸ್ಥಳದಲ್ಲಿ ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯ ಮೇಲೆ ಸಹಿ ಮಾಡಬೇಕು.
8. ಅಭ್ಯರ್ಥಿಯು ಪರೀಕ್ಷೆಯ ನಂತರ ಕೊಠಡಿ ಮೇಲ್ವಿಚಾರಕರಿಗೆ ಮೂಲ ಓ.ಎಂ.ಆರ್. ಉತ್ತರ ಹಾಳೆ ಮತ್ತು ವಿಶ್ವವಿದ್ಯಾನಿಲಯದ ಪ್ರತಿಯನ್ನು ಹಿಂದಿರುಗಿಸಬೇಕು.
9. ಅಭ್ಯರ್ಥಿಯು ಪ್ರಶ್ನೆ ಪುಸ್ತಕವನ್ನು ಮತ್ತು ಓ.ಎಂ.ಆರ್. ಅಭ್ಯರ್ಥಿಯ ಪ್ರತಿಯನ್ನು ತಮ್ಮ ಜೊತೆ ತೆಗೆದುಕೊಂಡು ಹೋಗಬಹುದು.
10. ಕ್ಯಾಲ್ಕುಲೇಟರ್, ಪೇಜರ್ ಮತ್ತು ಮೊಬೈಲ್ ಫೋನ್‌ಗಳನ್ನು ಪರೀಕ್ಷಾ ಕೊಠಡಿಯ ಒಳಗೆ ಅನುಮತಿಸಲಾಗುವುದಿಲ್ಲ.
11. ಅಭ್ಯರ್ಥಿಯು ದುಷ್ಕೃತ್ಯದಲ್ಲಿ ತೊಡಗಿರುವುದು ಕಂಡುಬಂದರೆ, ಅಂತಹ ಅಭ್ಯರ್ಥಿಯನ್ನು ಕೋರ್ಸ್‌ಗೆ ಪರಿಗಣಿಸಲಾಗುವುದಿಲ್ಲ ಮತ್ತು ನಿಯಮಗಳ ಪ್ರಕಾರ ಇಂತಹ ಅಭ್ಯರ್ಥಿಯ ವಿರುದ್ಧ ಕ್ರಮ ಕೈಗೊಳ್ಳಲಾಗುವುದು.
ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯನ್ನು ತುಂಬಲು ಸೂಚನೆಗಳು
1. ಪ್ರತಿಯೊಂದು ಪ್ರಶ್ನೆಗೆ ಒಂದೇ ಒಂದು ಅತ್ಯಂತ ಸೂಕ್ತವಾದ/ಸರಿಯಾದ ಉತ್ತರವಿರುತ್ತದೆ.
2. ಪ್ರತಿ ಪ್ರಶ್ನೆಗೆ ಒಂದು ವೃತ್ತವನ್ನು ಮಾತ್ರ ನೀಲಿ ಅಥವಾ ಕಪ್ಪು ಬಾಲ್ ಪಾಯಿಂಟ್ ಪೆನ್‌ನಿಂದ ಮಾತ್ರ ತುಂಬತಕ್ಕದ್ದು. ಉತ್ತರವನ್ನು ಮಾರ್ಪಡಿಸಲು ಪ್ರಯತ್ನಿಸಬೇಡಿ.
3. ವೃತ್ತದೊಳಗಿರುವ ಅಕ್ಷರವು ಕಾಣದಿರುವಂತೆ ವೃತ್ತವನ್ನು ಸಂಪೂರ್ಣವಾಗಿ ತುಂಬುವುದು.
4. ಓ.ಎಂ.ಆರ್. ಹಾಳೆಯಲ್ಲಿ ಯಾವುದೇ ಅನಾವಶ್ಯಕ ಗುರುತುಗಳನ್ನು ಮಾಡಬೇಡಿ.